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9.) Using The Attached Table, Estimate The Change In Energy For The Following

Reaction. $\text{CH}_2\text{CH}_2(\text{g}) + \text{Br}_2(\text{l}) \rightarrow 2 \text{BrCH}_2\text{Br}(\text{l}) - 92 \text{ KJ/mol}$ 10.) Given The Following Information Calculate The Lattice Enthalpy Of Calcium Chloride. Page 6 Of 7 1th, 2024

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Equation: $2 \text{K} + \text{Cl}_2 \rightarrow 2 \text{KCl}$ How Many Grams Of KCl Is Produced From 2.50 G Of
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