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Near The Oxygen End Of The Molecule And Partial
Positive Charge Near The Hydrogen Atoms. Thus, The
Negative End Of A Water Molecule Is Strongly
Attracted To ... 4th, 2024 Chapter 3

Stoichiometry Chapter 3 Stoichiometry 3-7 EXAMPLE
PROBLEM: Determine The Molar Mass Of A Compound
Calculate The Molar Mass For Each Of The Following
Compounds: (a) 2-Propanol, $\text{CH}_3\text{CH}(\text{OH})\text{CH}_3$ (b) Iron(II)
Phosphate, $\text{Fe}_3(\text{PO}_4)_2$ SOLUTION: You Are Asked To
Calculate The Molar Mass Of A Compound. You Are
Given The Compound's Formula. 3th, 2024.

CHAPTER 2: Atoms, Molecules And Stoichiometry 2.5
Stoichiometry And Equations Stoichiometry 1)
Stoichiometry Is The Proportion Of Things Either
Reacting Or Combining. 2) In Compounds, It Refers To
The Ratio In Which The Atoms Are Combined Together.
For Example, Water, H_2O Has A Stoichiometry Of 2
Hydrogen To 1 Oxygen. 3) It Also Refers To The
Reacting Proportions In A Chemical Equation. 2th,
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