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Answers Nov 08, 2021 · Chapter 12 1 Stoichiometry Worksheet Chapter 12 1 Stoichiometry Worksheet Read PDF Chapter 12 1 Stoichiometry Worksheet Answers Figure 11.4.1 Is Shown In Figure 12.2.1. We Can Use The Balanced Chemical Equation For The Reaction And Either The Masses Of Solid Reactants And Products Mar 1th, 2024.

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33. $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$
Actual yield = 1.45g Theoretical yield = 5.95g
 $\text{KClO}_4 \times 1 \text{ mol } 122.55 \text{ g} \times 2 \times 16 = 32.00 \text{ g}$
 $1 \text{ mol } = 3.11 \text{ g O}_2$
% Yield = $\frac{1.45 \text{ g } 3.11 \text{ g}}{3.11 \text{ g}} \times 100\% = 46.6\%$
34. $4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$
Percent purity = 62.8%
pure FeCO_3 May 1th, 2024
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