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Name: KEY

Chemistry 11 Stoichiometry Review Worksheet

Name: KEY 33. $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$

Actual yield = 1.45g Theoretical yield = 5.95g $\text{KClO}_4 \times 1\text{mol}$

$122.55\text{g} \times 2\text{O}_2 \times 32.00\text{g} \times 1\text{mol} = 3.11\text{g O}_2$

%Yield = $1.45\text{g} \times 3.11\text{g} \times 100\% = 46.6\%$ 34. $4\text{FeCO}_3 + \text{O}_2$

$\rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$

Purity = 62.8%? pure FeCO_3 May 1th, 2024

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R. Janssen ...

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Chapter 4 Aqueous Reactions And Solution
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11.4.1 Is Shown In Figure 12.2.1. We Can Use The
Balanced Chemical Equation For The Reaction And

Either The Masses Of Solid Reactants And Products Jun 3th, 2024

AP Chemistry: Stoichiometry - Multiple Choice Answers

AP Chemistry: Stoichiometry - Multiple Choice Answers

44. What Number Of Moles Of O_2 Is Needed To Produce 14.2 Grams Of P_4O_{10} From P? (Molar Mass $P_4O_{10} = 284$) (A) 0.0500 Mole (B) 0.0625 Mole (C) 0.125 Mole Jan 3th, 2024

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