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If Pressure Is Needed In kPa Then Convert By Multiplying By $101.3 \text{ kPa} / 1 \text{ atm}$ to Get. $R = 8.31 \text{ L}\cdot\text{kPa} / (\text{K}\cdot\text{mole})$ | Feb 3th, 2024

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Worksheet 7 - Ideal Gas Law I. Ideal Gas Law The Findings Of 19th Century Chemists And Physicists,

Among Them Avogadro, Gay-Lussac, Boyle And Charles, Are Summarized In The Ideal Gas Law: $PV = nRT$ P = Pressure V = Volume n= Moles Of Gas, R = Universal Gas Constant T = Temperature. The Value Of R Varies With The Jul 3th, 2024

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A.) _____ Write The Combined Gas Law. B.) _____ At What Kelvin Temperature Is There No Kinetic Energy? C.) Match The Graph Below With The Law: Boyle's Or Charles Boyle's Or Charles 2. Solve The Following Problems. Show Your Wo May 3th, 2024

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