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$N_2O + ne \rightleftharpoons N_2$ (2.1.36) The Free Energy Of The Cell Reaction Is Given By, $\Delta G =$
 $\Delta G^0 + RT \ln \frac{a_{N_2}}{a_{N_2O}}$ (2.1.38) Which Since $\Delta G = -nFE_{rxn}$
And $a_{H^+} = a_{H^+}^0 = 1$ Becomes, $E = E^0 + \frac{RT}{nF} \ln \frac{a_{N_2}}{a_{N_2O}}$. (2.1.40) This

Expression Is Called The Nernst Equation. This Relation Can 2th, 2024

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