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$\text{N}^{\text{OO}} + \text{ne}^- \rightleftharpoons \text{N}^{\text{RR}}$  (2.1.36) The Free Energy Of The Cell Reaction Is Given By,  $\Delta G =$   
 $\Delta G^0 + RT \ln \frac{a_{\text{R}}}{a_{\text{O}}} \frac{a_{\text{H}^+}}{a_{\text{H}_2}}$  (2.1.38) Which Since  $\Delta G = -nF E_{\text{rxn}}$   
 And  $\frac{a_{\text{H}^+}}{a_{\text{H}_2}} = 1$  Becomes,  $E = E^0 + \frac{RT}{nF} \ln \frac{a_{\text{R}}}{a_{\text{O}}}$ . (2.1.40) This

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$$\text{Ru N N N N N N N N N N N N } 2 + \text{N N N N N N } 2 \text{ Ru N N N N N } 2 + \text{Ru N N N N N N } 2 +$$
$$\text{Ir N N F F F F Ir N N N N F F F F CF } 3 \text{ F } 3 \text{ C TBu TBu } 3 + 3 + 1 + \text{Ir N N N TBu TBu } 3 + \text{Ir}$$
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Zhang Et Al., Matter4, 164-181 1th, 2024

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4 1 Introduction To Electrochemical Cells Table 1.1 Dimensions Of Commercially Available Battery Sizes [1] . Battery Size Diameter (mm) Height (mm) N 12 30.2 AAA 10.5 44.5 AA 14.5 50.5 C 26.2 50 D 34.2 61.5 F 32.0 91.0 Flat Cells 1th, 2024

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Electrochemical Cells Revised 12/8/14 2 (2) E Cell =  $E^\circ$  Cell -  $(0.059 / N) \log Q$  (Nernst Equation) Q: Reaction Quotient N: Moles Of E – Transferred For The General Cell Reaction (3) Occurring At 298 K, The Nernst Equation Can Be Rewritten As 1th,

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