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$N_2 + 2H^+ + 2e^- \rightarrow N_2H_4$ (2.1.36) The Free Energy Of The Cell Reaction Is Given By, $\Delta G = \Delta G^0 + RT \ln \frac{a_{N_2} a_{H^+}^2}{a_{N_2H_4}}$ (2.1.38) Which Since $\Delta G = -nFE_{rxn}$ And $a_{H^+} = 1$ Becomes, $E = E^0 + \frac{RT}{nF} \ln \frac{a_{N_2}}{a_{N_2H_4}}$ (2.1.40) This Expression Is Called The Nernst Equation. This Relation Can 3th, 2024

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Elena Arroyo-de Dompablo, ZhaopingLiu,YingShirley Meng Liuzp@nimte.ac.cn (Z.L.) Shmeng@ucsd.edu (Y.S.M.) HIGHLIGHTS
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