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The yield = $5.95 \text{g KClO}_4 \times \frac{1 \text{ mol}}{122.55 \text{g}} \times 2 \times \frac{1 \text{ mol}}{32.00 \text{g}} = 3.11 \text{g O}_2$
% Yield = $\frac{1.45 \text{g}}{3.11 \text{g}} \times 100\% = 46.6\%$
 $4 \text{FeCO}_3 + \text{O}_2 \rightarrow 2 \text{Fe}_2\text{O}_3 + 4 \text{CO}_2$ Percent purity = 62.8%? g pure FeCO₃ 2th, 2024

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