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Mole Fraction Sugar Of Solution  $\chi$  Sugar = (0.292 Mol)/(0.292 + 55.6) = .00522 Mole Fraction Water Of Solution  $\chi$  Water = (55.6 Mol)/(0.292 + 55.6) = 0.995

(or 1-  $\chi$  Sugar) Molality M = (.292 Mol)/(1 Kg) = 0.292 Mol Kg-1 Molarity M = (.292 Mol)/(1 3th, 2024

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3.mole Fraction = 0.950 4.mole Fraction = 0.0360 5.mole Fraction = 0.925 Nsci140802exam 018 10.0points 3.3 G Of NaCl (58.4 G/mol) And 9 G Of KBr (119.8 G/mol) Were Dissolved In 40 G Of Water (18.0 G/mol). What Is The Mole Fraction Of KBr In The Solution? Nsci140805 019 10.0points How Many G 2th, 2024

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As 1 Atm =  $1.013 \times 105$  Pa, The Osmotic Pressure Is  $\Pi$  =  $(0.0036 \times 1.013 \times 105)$  Pa = 365 Pa Hence C =  $\Pi$  / RT = (365 Pa) / ((8.314 M3 Pa K-1 Mol-1)  $\times$  (293 K)) = 0.15 Mol M-3 As 1 M3 = 1000 L, This Corresponds To C =  $1.5 \times 10$  1th, 2024

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