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What Would Be The Weight Of This Maximum Yield? To Answer This, We Need To Convert Our Mole Value For H_2S To A Mass Value. Remember That $M = nM_w$ We Can Specify $M_w = (2 \times 1,008) + 32.06 = 34.08 \text{ g/mol}$
 $M = 3.5 \times 34.08 = 119.3 \text{ grams}$
 H_2S could be produced. III. If You Would Like To Produce 3.00 kg H_2S
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Model 2: The Mole Concept 1 Dozen Items = 12 Dozen
Items = 1 Egg Sandwich = 1 Mole = Key Questions C
12 Items 1 Gross Of Items 3 Eggs And 2 Slices Of
Bread 6.02×10^{23} items (Avogadro's Number) Use
Dimensional Analysis ONLY For All Calculations. Show
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$\times 10^{23}$ molecules) 1 Mole 1 Mole Pogil Relative Mass And The Mole Answers Tip Video. Chemistry Questions And Answers.) So It May Need Some Further Explanation. Where Do You Find The Molar Mass Of An

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